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Open

Comparative Analysis Worksheet

What Is – What Isn't Information

What happened?	What didn't happen?
What happened?	What might you have expected to happen, but didn't?
When does it occur?	When doesn't it occur?
Who has it happened to?	Who hasn't it happened to?
Where did it happen?	Where didn't it happen?
What changed in the process?	What didn't change in the process?
Which supplier(s) was involved?	Which supplier(s) was not involved?
Which customer(s) was involved?	Which customers were not involved?
What changes were made to the process recently?	What wasn't changed in the process recently?
What measurement system was used?	What measurement system wasn't used?



Date _____

Neutralization Reactions Worksheet

1. Write the balanced chemical equations for the neutralization reactions between

- a) HI and NaOH
 $\text{HI} + \text{NaOH} \rightarrow \text{H}_2\text{O} + \text{NaI}$
- b) H_2CO_3 and $\text{Sr}(\text{OH})_2$
 $\text{H}_2\text{CO}_3 + \text{Sr}(\text{OH})_2 \rightarrow 2 \text{H}_2\text{O} + \text{SrCO}_3$
- c) $\text{Ca}(\text{OH})_2$ and H_3PO_4
 $3 \text{Ca}(\text{OH})_2 + 2 \text{H}_3\text{PO}_4 \rightarrow 6 \text{H}_2\text{O} + \text{Ca}_3(\text{PO}_4)_2$
- d) hydrobromic acid and barium hydroxide
 $2 \text{HBr} + \text{Ba}(\text{OH})_2 \rightarrow 2 \text{H}_2\text{O} + \text{BaBr}_2$
- e) zinc hydroxide and nitric acid
 $\text{Zn}(\text{OH})_2 + 2 \text{HNO}_3 \rightarrow 2 \text{H}_2\text{O} + \text{Zn}(\text{NO}_3)_2$
- f) aluminum hydroxide and hydrochloric acid
 $\text{Al}(\text{OH})_3 + 3 \text{HCl} \rightarrow 3 \text{H}_2\text{O} + \text{AlCl}_3$

2. Complete and balance the following equations representing neutralization reactions:

- a) $2 \text{CsOH} + \text{H}_2\text{CO}_3 \rightarrow 2 \text{H}_2\text{O} + \text{Cs}_2\text{CO}_3$
- b) $2 \text{HF} + \text{Mg}(\text{OH})_2 \rightarrow 2 \text{H}_2\text{O} + \text{MgF}_2$
- c) $3 \text{HNO}_3 + \text{Al}(\text{OH})_3 \rightarrow 3 \text{H}_2\text{O} + \text{Al}(\text{NO}_3)_3$
- d) $\text{HCl} + \text{KOH} \rightarrow \text{H}_2\text{O} + \text{KCl}$
- e) $\text{HBrO}_3 + \text{LiOH} \rightarrow \text{H}_2\text{O} + \text{LiBrO}_3$



